

# Chem151-Tutorial Assignment #1 (Fall 2009)

## What's in a Name?

### The Nomenclature of Inorganic Compounds

Name Master  
Last first

Chem ID \_\_\_\_\_

TA \_\_\_\_\_

Neatly write the answers to the following questions. Submit the document to the Green locker in the 3<sup>rd</sup> floor hallway of Laboratory Science Building before 5 pm 9/14/09 (Monday).

(16)1. Fill in the blanks with the names or formulas of the compounds/ ions.

Formula	Name	Formula	Name
$\text{Cu}_2\text{O}$	<i>Copper(I) oxide</i>	$\text{Mg}(\text{CH}_3\text{COO})_2$	Magnesium acetate
$\text{Ba}(\text{MnO}_4)_2$	<i>Barium permanganate</i>	$\text{HI}(\text{aq})$	Hydroiodic acid
$\text{GaAsO}_4$	<i>Gallium arsenate</i>	$\text{Fe}(\text{HSO}_4)_3$	Ferric hydrogen sulfate
$\text{CsBrO}_4$	<i>Cesium perbromate</i>	$\text{Hg}_2\text{Cl}_2$	Mercury(I) chloride
$\text{H}_2\text{S}(\text{aq})$	<i>Hydrosulfuric acid</i>	$\text{Al}(\text{HCO}_3)_3$	Aluminum bicarbonate
$\text{CCl}_4$	<i>Carbon tetrachloride</i>	$\text{K}_2\text{C}_2\text{O}_4$	Potassium oxalate
$\text{Na}_2\text{O}_2$	<i>Sodium peroxide</i>	$\text{HBrO}_4$	Perbromic acid
$\text{OF}_2$	<i>Oxygen difluoride</i>	$\text{XeF}_4$	Xenon tetrafluoride

(4)2. Based on the Table of Common Polyatomic Ions (in the **Related Tutorial: "Molecular Representations"**), circle the smallest anion in each group in the following:

- a)  $\text{CN}^-$   $\text{OH}^-$   $\text{NO}^-$   $\text{CH}^-$
- b)  $\text{ClO}_4^-$   $\text{MnO}_4^-$   $\text{SO}_4^{2-}$   $\text{PO}_4^{3-}$
- c)  $\text{NO}_3^-$   $\text{CO}_3^{2-}$   $\text{SO}_3^{2-}$   $\text{PO}_3^{3-}$